

## Carbon Nano-Painting: Application to non-Phosphate Oxyanions, e.g. Borates

A. Abouimrane, M. Armand, N. Ravet

CNRS-UdM international Laboratory UMR 2289,  
Département de Chimie, Université de Montréal  
P.O. Box 6128 Montréal QC Canada

Boron compounds have found little use in rechargeable lithium batteries other than as dopants of lamellar oxides.  $\text{FeBO}_3$  ( $\bar{V} \approx 1.6$  V vs.  $\text{Li}^\circ/\text{Li}^+$ ) was suggested as an anode [1,2] and  $\text{LiFeBO}_3$  ( $\bar{V} \approx 2.9$  V) was used as a cathode [3]. This study [3] showed that  $\text{LiFeBO}_3$  has a capacity of less than 5% of theoretical (220 mA/g) even on the first cycle.  $\text{LiFeBO}_3$  crystallizes in the monoclinic system  $\text{C}2/c$  [4]. In the environment of iron atoms, the five oxygens form a trigonal-bipyramidal coordination shell.

In this study we report a new route for the synthesis of electronically (carbon coated) conductive  $\text{LiFeBO}_3$  particles. In a typical preparation, appropriate amounts of  $\text{FeC}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$ ,  $\text{LiBO}_2$  and cellulose ester were mixed and then fired at 700 °C for 24 h under argon. The carbon represents less than 2%, by weight, of the final product. The crystallographic structure has been confirmed by X-ray powder diffraction.

Electrochemical performances were evaluated at 80 °C using a coin cell and a polymer electrolyte. Cyclic voltammetry and galvanostatic cycling curves of cells with either  $\text{LiFeBO}_3$ , with/without carbon, are shown in Fig. 1 and 2. A capacity of about 156 mA/g at an average voltage of 2.9 V vs.  $\text{Li}^\circ/\text{Li}^+$  was obtained at the first discharge (lithium de-insertion) and was stable on cycling.

The nano-painting concept appears here to be applicable to a variety of oxides, whose low electronic conductivity had up to now precluded their use as cathode materials.

$\text{LiFeBO}_3$ , with its large theoretical capacity, has an energy content roughly similar to that of  $\text{LiFePO}_4$ . This study is only preliminary, and progress is likely to be reported for the use of this promising cathode material in rechargeable lithium batteries.

### References

- [1] J. L. C. Rowsell, J. Gaubicher and L. F. Nazar *J. Power Sources*, 97-98 (2001), 254-257.
- [2] A.I. Palos, M. Morcette and P. Strobel *J. Solid State Electrochem.* **6** (2002), 134-138.
- [3] V. Legagneur, Y. An, A. Mosbah, R. Portal, A. Le Gal La Salle, A. Verbaere, D. Guyomard and Y. Piffard, *Solid State Ionics*, **139**, (2001), 37-46.
- [4] O. S. Bondareva, M. A. Simonov, Yu. K. Egorov-Tismenko, and N. V. Belov, *Sov. Phys. crystallogr.* **23**, (1978), 269-271.

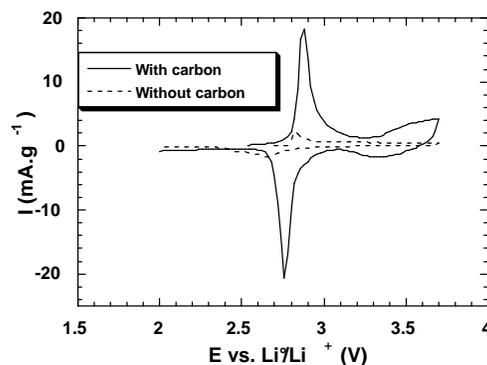


Fig. 1 : Cyclic Voltammetry of  $\text{LiFeBO}_3$  at a sweep rate of 20 mV/h in the potential range 2 V  $\leftrightarrow$  3.7 V

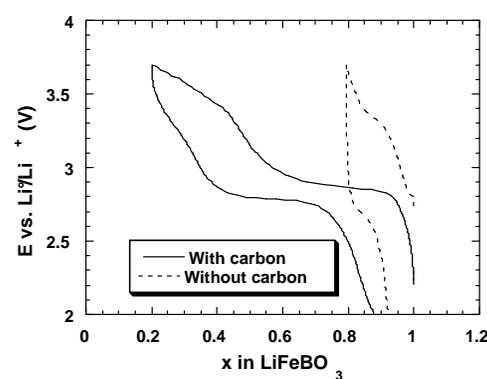


Fig. 2 : Charge-discharge curves for  $\text{LiFeBO}_3$  at C/12 in the potential range 2 V  $\leftrightarrow$  3.7 V